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vaporization

evaporation Most of
the molecules do not
have enough kinetic
energy to overcome
the attractive forces.

As the temperature is
increased, the average
kinetic energy

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increases and more particles have enough kinetic energy to overcome the forces keeping them in the liquid state.

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STATES OF MATTER
13

b. The minimum temperature at which the substance exists in the liquid state. c. The maximum pressure and temperature at which the substance is

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in equilibrium between
the solid, liquid, and
gas. d. The pressure
and temperature at
which the substance is
in equilibrium between
the solid, liquid, and
gas states.

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Terms in this set (35)
kinetic energy. The
energy an object has
because of its motion.
kinetic theory. explains

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Answers
the states of matter,
based on the concept
that all matter consists
of tiny particles that
are in constant motion.
The particles in a gas
are considered to be
small, hard spheres
with an insignificant
volume.

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13.1 The Fluid States

300 States of Matter

FIGURE 13-1 The ice

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cube, a solid, has a definite shape. But water, a fluid, takes the shape of its container.

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Gas particles are a) free and fast moving
b) less free and slower moving
c) tightly packed and can't move
d) free and slow moving, Elastic collisions
a) loses

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energy and the total kinetic energy changesb) no loss but change in total kinetic energyc) no loss and no changes in total kinetic energyd) loses of energy and no loss in total kinetic energy, Gas pressurea) force taken away from collisionsb ...

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Gas particles area) free

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Answers

- and fast movingb) less free and slower movingc) tightly packed and can't moved) free and slow moving, Elastic collisionsa) loses energy and the total kinetic energy changesb) no loss but change in total kinetic energyc) no loss and no changes in total kinetic energyd) loses of energy and no loss in total kinetic energy, Gas pressurea) force

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taken away from
collisionsb...

States of Matter: Chapter 13

Feb 109:15 AM.

Chapter 13 "States of
Matter". Feb 109:15
AM. •OBJECTIVES:

- Describe the
assumptions of the
"kinetic theory" as it
applies to gases. Feb
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- OBJECTIVES:

- Interpret gas pressure
in terms of kinetic

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theory.
Answers

**Chapter 13 “States
of Matter”**

chemistry vocabulary
chapter 13 states
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energy an object has
because of its motion.
An empty space with
no particles and no
pressure. the energy
an object has because
of its motion.

**chemistry
vocabulary chapter**

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and Solids In your
textbook, read about
liquids and solids. In
the space at the left,
write true if the

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statement is true; if the
statement is false,

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Chapter 10 - Chemical
Reactions; Chapter 11 -
The Mole; Chapter 12 -

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Chapter 14 - Gases;
Chapter 15 - Solutions;
Chapter 16 - Energy
and Chemical Change;
Chapter 17 ...

Chapter 13 - States of Matter - Ms. K Kelly - John F ...

An empty space with
no particles and no
pressure., A measure
of the force exerted by
a gas above a liquid in
a sealed container; a

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dynamic equilibrium exist between the vapor and the liquid., A solid in which the atom, ions, or molecules are arranged in an orderly, repeating, three-dimensional pattern called a _____ lattice., The process in which a solid changes to a gas or vapor without ...

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There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are solid, liquid, and gas. These three states are quite different. The main difference is in their particles.

Chapter 13: States of Matter -

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the process in which molecules of a liquid escape from the surface of a liquid that is not boiling. boiling point. temperature at which the vapor pressure of a liquid is just equal to the external pressure. crystal. sample in which particles are arranged in an orderly, repeating, three-dimensional pattern.

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Kinetic energy. the
energy an object has
because of its motion.
Kinetic theory. All
matter consists of tiny
particles that are in
constant motion ...

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Chapter 13 States of Matter Study Guide - StudyBlue

Chapter 13 States of Matter pages 341 to 362. Properties of fluids. Gases and liquids are both fluids. Both these states of matter have greater freedom of motion. Objects exert pressure. Pressure is the magnitude of the force divided by the area of contact. The unit of

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measure for pressure is
the Pascal (Pa) 1
newton/ square meter.

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