

## Chapter 12 Stoichiometry Vocabulary Review Answers

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Chapter 12 Review Stoichiometry Important Vocabulary • Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry • Mole ratio- a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles

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Chapter 12 REVIEW: Stoichiometry, Theoretical, Actual & Percent yield Part I. Stoichiometry 1.  $1 \text{ ZnI}_2 \rightarrow 1 \text{ Zn} + 1 \text{ I}_2$  How many grams of iodine will you produce if you begin your rxn with 56.7g of zinc iodide?

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STOICHIOMETRY 12 Vocabulary Review Match the correct vocabulary term to each numbered statement. Write the letter of the correct term on the line. Column A 1. the starting materials in a chemical reaction 2. a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles

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Chapter 12 (pages 320 - 321) - Answer questions - 1-9, 14. End-of-Chemistry Assessment Review. To help you prepare for the End-of-Chemistry Assessment, review the following at the end of each chapter: Chapter Highlights. Vocabulary . Answer Chapter Review questions (answers will be posted on my website next week)

### **End-of-Chemistry Assessment Review**

CHAPTER 12 378 Chapter 12 Study Guide Study Tip Prioritize Schedule your time realistically. Stick to your deadlines. ... Vocabulary Review, Quiz, Chapter Test A, Chapter Test B ... Stoichiometry 379 CHAPTER 12 Assessment 36. a. Two formula units  $\text{KClO}_3$  decompose to form two formula units  $\text{KCl}$

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Chapter 11 Small-Scale Lab Section 11.3 Precipitation Reactions: Formation of Solids, page 345 Analysis 1. a. Na<sub>2</sub>CO<sub>3</sub> ... Section Review 12.1 Part A Completion moles/molecules 2. balanced equation 3. mass/atoms 4. atoms/mass ... Vocabulary Review 12 1. d 5. g 8. i 2. h 6. j 9. e 3. f 7. b 10. c 4. a Quiz for Chapter 12 1. coefficients 6. NT 2 ...

### **Chapter 11 Small-Scale Lab**

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## Chapter 12 Stoichiometry Answer Key

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: N<sub>2</sub>(g) + 3H<sub>2</sub>(g) ...

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